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Other Results for
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Chemical Equations

And Stoichiometry

Practice Problems

Answers: Practice

Problems:

Stoichiometry. Practice

Problems:

Stoichiometry. Balance

the following chemical

reactions: Hint a. $\text{CO} +$

$\text{O}_2 \rightarrow \text{CO}_2$ b. $\text{KNO}_3 \rightarrow \text{KNO}$

$2 + \text{O}_2$ c. $\text{O}_3 \rightarrow \text{O}_2$ d.

$\text{NH}_4\text{NO}_3 \rightarrow \text{N}_2\text{O} + \text{H}_2$

O e. $\text{CH}_3\text{NH}_2 + \text{O}_2$

$\text{CO}_2 + \text{H}_2\text{O} + \text{N}_2$

Hint f.

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Chemical Equations And Stoichiometry Practice Problems Answers

Honors Chemistry

Extra Stoichiometry

Problems 1. Silver nitrate reacts with barium chloride to form silver chloride and

barium nitrate. a. Write and balance the

chemical equation. 2

$\text{AgNO}_3 + \text{BaCl}_2 \rightarrow$

$\text{AgCl} + \text{Ba}(\text{NO}_3)_2$ b. If

39.02 grams of barium chloride are reacted in

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an excess of silver
nitrate, how many

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AP Chemistry:

Stoichiometry -

Multiple Choice

Answers 44. What

number of moles of O_2

is needed to produce

14.2 grams of P_4O_{10}

from P? (Molar Mass P_4O_{10}

$= 284$) (A)

0.0500 mole (B) 0.0625

mole (C) 0.125 mole

(D) 0.250 mole (E)

0.500 mole $4 P + 5 O_2$

$\rightarrow P_4O_{10}$ >> 14.2 g P

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40 10 x 1 mol P 40 10

x 5 mol O 2 =

0.250 mol 284 g P 40

**AP Chemistry:
Stoichiometry -
Multiple Choice
Answers**

Play this game to review Chemistry. $N_2 + 3 H_2 \rightarrow 2 NH_3$ How many moles of H_2 are needed to react with 2 moles of N_2 ? ... $N_2 + 3 H_2 \rightarrow 2 NH_3$ How many moles of H_2 are needed to react with 2

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**Stoichiometry
questions (practice)**

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20 Then do some stoichiometry using "easy math" 16 g of methane (MM = 16) is 1 mole and 1 mole of methane will produce 1 mole of $\text{CO}_2 = 44 \text{ g}$, and 2 moles of H_2O which is 36 g for a total

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of 80 g 4. d Balance: C
 $3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 +$
4H₂O 5. d Balance:
 $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$

Practice Test Ch 3
Stoichiometry Name
Per

Stoichiometry Chapter
Exam Instructions.

Choose your answers
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